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Practice Problems

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What is the difference between an oxidation-reduction reaction and a half-reaction? 2.

What is the function of the salt bridge in an electrochemical cell? 3.

What is the criterion for spontaneous chemical change based on cell potentials? Explain. 4.

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Electrochemistry Practice Problems

Electrochemistry

Problems 1) Given the E° for the following half-reactions: $\text{Cu}^+ + e^- \rightleftharpoons \text{Cu}^\circ$ $E^\circ_{\text{red}} = 0.52 \text{ V}$

$\text{Cu}^{2+} + 2e^- \rightleftharpoons \text{Cu}^\circ$ $E^\circ_{\text{red}} = 0.34 \text{ V}$ What is E° for the reaction: $\text{Cu}^+ \rightleftharpoons \text{Cu}^{2+} + e^-$

2) How many Faradays are required to produce 21.58 g of silver from a silver nitrate solution?

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Solutions for
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Problem Set Constants:
F 96484.56.coul .mole
1 T (273.15 25) K M
mole R
8.31441.joulemole liter
1.K 1 Equations E
std_cell E cathode E
anode E cell E std_cell
R.T n.F In C anode C
cathode. 1 a. Calculate
the cell potential and

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free energy available
for the following
electrochemical
systems

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Practice Problems. 1.
An atom with the
electron configuration
 $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^2$ has an
incomplete. 2p
sublevel. Second
principal energy level.

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Third principal energy level.

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Solution: (a) The reduction reaction is.

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$\text{Al}^{3+} + 3\text{e}^- \rightarrow \text{Al}$. Thus, 3 mole of electrons are needed to reduce 1 mole of Al^{3+} . $Q = 3 \times F = 3 \times 96500 = 289500$ coulomb. (b)

The reduction is.



$\text{Mn}^{2+} + 4\text{H}_2\text{O}$. 1 mole 5 mole. $Q = 5 \times F = 5 \times 96500 = 48500$ coulomb.

Solved Examples On Electrochemistry - Study Material for ...

6. Answer the following

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questions about
Answers
electrochemistry. (a)
Several different
electrochemical cells
can be constructed
using the materials
shown below. Write the
balanced net-ionic
equation for the
reaction that occurs in
the cell that would
have the greatest
positive value of E_{cell} .
 $\text{Al(s)} \rightarrow \text{Al}^{3+}(\text{aq}) + 3$
 e^- $\text{Cu}^{2+}(\text{aq}) + 2 \text{e}^- \rightarrow$
 Cu(s)

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Electrochemistry Free Response Questions

electrochemistry to the
thermodynamic

concept of work, free
energy, through the
equation: free energy
 $= \Delta G = -q E = -nFE$

You will also remember
that free energy $= \Delta G$
 $= -RT \ln K$ From this
equation, the following
must be true about
spontaneous reactions:

type of reaction

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thermodynamics
electrochemistry
equilibria spontaneous
reaction

Chapter 21:

ELECTROCHEMISTRY

TYING IT ALL

TOGETHER

If it displaces Au + (aq) from solution, then it has a reduction potential smaller than $E^\circ_{\text{Au}^+/\text{Au}}$ = 1.68V. But if it does not displace Fe³⁺ + (aq) from solution, then its

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reduction potential is larger than. $E^\circ \text{Fe}^{3+} + (aq) / \text{Fe}^{2+} + (s) = 0.769\text{V}$. Therefore, $0\text{V} < E^\circ < 0.17\text{V}$.

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AP REVIEW QUESTIONS

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Answers Answer: (a) tin

electrode is the

cathode; cathode is the

site of reduction (gain

in electrons) and will

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convert metal ions into
a metal. (b) (see
diagram) (c) red: Sn^{2+}
(aq) + 2 e⁻ Sn (s) $E^\circ =$
-0.14 V oxid: X (s) - 3
e⁻ X³⁺ (aq) $E^\circ =$
+0.74 V $E^\circ_{\text{cell}} =$
+0.60 V red: X³⁺ (aq)
+ 3 e⁻ X

AP REVIEW

QUESTIONS

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Answers

If you are stumped,
answers to numeric
problems can be found

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Answers

by clicking on "Show Solution" to the right of the question. Do NOT type units into the answer boxes, type only the numeric values. Do NOT use commas or scientific notation when entering large numbers. Answer all non-integer questions to at least 3 significant figures.

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Exercises**

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QUESTIONS SOLVED.

3.1. How would you determine the standard electrode potential of the system $\text{Mg}^{2+} + 1 \text{Mg}$? Ans: A cell will be set up consisting of $\text{Mg}/\text{MgSO}_4 (1 \text{ M})$ as one electrode and standard hydrogen electrode $\text{Pt}, \text{H}_2 (1 \text{ atm})/\text{H}^+ (1 \text{ M})$ as second electrode, measure the EMF of the cell and also note the direction of deflection in the

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voltmeter.

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Electrochemistry is the branch of physical chemistry which deals with the study of the relationship between electricity, as a measurable and quantitative phenomenon, and identifiable chemical change, with either electricity, considered

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an outcome of a particular chemical change or vice versa. Electrochemistry MCQs. 1.

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at the earliest.

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Chapter 3
Electrochemistry**

Answer. Oxidation-reductions reactions always have an electron transfer from the oxidized species to the reduced species. When the oxidized species is separated from the reduced species, a balanced reaction can be written

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for each process (oxidation or reduction) that is called a half-reaction. All half-reactions must have electrons either as reactants (for reduction half-reactions) or products ...

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meaning of chapter in
the best manner. The
products formed at
either electrode is
given in terms of
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